

# Pensar La Biblia

**Terry C. Jones**

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temperature and pressure which is equal to 22.4 l for 1 mole of any ideal gas at a temperature equal to 273.15 K and a pressure

#### **1.4 Avogadro's number and the mole chemistry**

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web the molar volume of a gas from Avogadro's law the volume is directly proportional to the number of moles. The important point to keep in mind is that  $n$  can be moles of any gas since one mole always contains  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy. Based on this principle it is found *molar volume of gases o level chemistry notes chem not* - Apr 30 2022

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web mar 12 2018 you can define the concentration of a gas. For example, if you have  $x$  moles of a gas in a container of volume  $V$ , the concentration is  $c = \frac{x}{V}$  with units of mol/dm<sup>3</sup>.

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web the molar volume has the SI unit of cubic metres per mole (m<sup>3</sup>/mol). Although it is more typical to use the units cubic decimetres per mole (dm<sup>3</sup>/mol) for gases and cubic centimetres per mole (cm<sup>3</sup>/mol) for liquids and solids.

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web zero degrees Celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure. The volume occupied by one mole of a gas is called molar volume. Therefore, the molar volume of a gas at standard temperature and pressure (STP) is 22.4 l.

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